

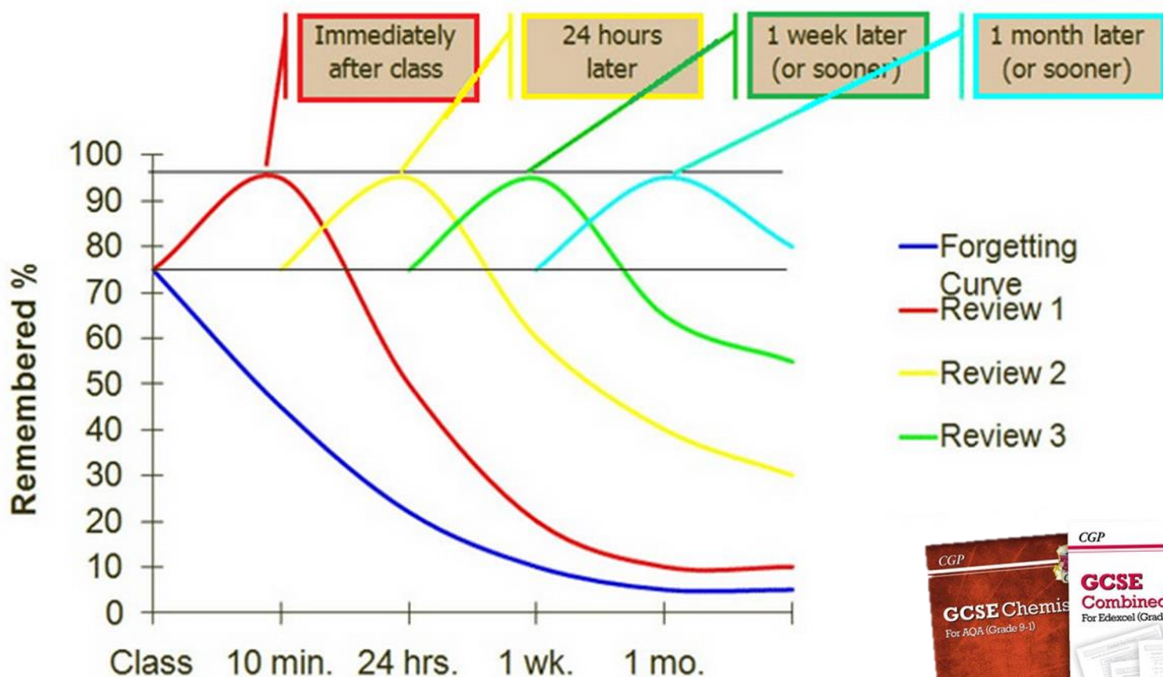
Regular review of information



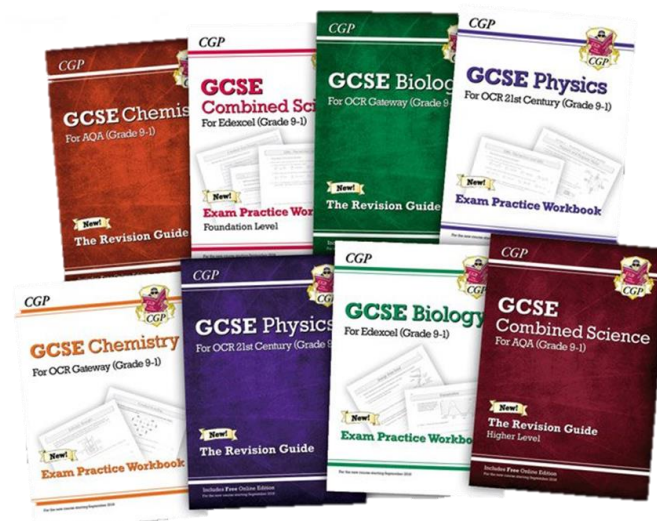
Holderness Academy
& Sixth Form College



THE CONSORTIUM
ACADEMY TRUST



— Forgetting Curve
— Review 1
— Review 2
— Review 3



How to beat the curve of forgetting

Use your revision guide & re-read the topics covered in class

- 24 hr review
- 7 day review
- 1 month review

To make it even more effective – make it **'ACTIVE'** and simplify that section

Covalent Bonding

These molecules might be simple, but you've still got to know about them. I know, the world is a cruel place.

Learn These Examples of Simple Molecular Substances

1) A size
2) Atom
3) Simple
4) You can

They share a pair of electrons. They can cause molecules to form. outer shell of electrons.

simple molecular substances. Only the electrons in the outer shell between two circles are shared. For example:

Covalent bond is a shared pair of electrons

Full outer shell has 2 or 8 electrons

Low MP because of weak intermolecular forces

Poor conductor because no free electrons

• Sometimes two pairs of electrons are shared between atoms in a molecule. This is called a double covalent bond. Both atoms gain two extra electrons. E.g. oxygen and carbon dioxide both have double covalent bonds.

• The structure of hydrogen chloride, methane and water. There can be more than one covalent bond in a compound. E.g. carbon forms four covalent bonds in methane and oxygen forms two covalent bonds in water.

Hydrogen

Methane, CH₄

• Single molecules aren't much bigger than single atoms. On the melting point scale, most simple molecular substances are gases or liquids at room temperature. Molecular compounds don't conduct electricity. This is because they don't have electrons or ions that can move through the substance.

• Some single molecules are soluble in water and some aren't.

Aspirational, Resilient, Respectful, Kind